

RHEOLOGY MODIFIED COMPOSITIONS AND MODIFICATION AGENTS**TECHNICAL FIELD**

This invention relates to the field of rheology modification agents for use in water or water-based fluids, and more particularly to the preparation of compositions exhibiting stress-dependent fluidity and agents useful for preparing such compositions.

BACKGROUND ART

The use of rheology modification agents, frequently thickening agents, for aqueous and hydrophylic fluids has been common practice in a large number of industries. These fluids include, for example, oil field drilling fluids, metal-working fluids, mining fluids, fire control fluids, hydraulic fluids, water-based paints and coating fluids, stripping fluids, and the like. For each of these, and other, applications, the rheology modification agents serve very specific purposes tailored to the function for which the fluid is being employed. Among these purposes are pressure resistance; suspension of solids; adjustment of reaction time(s); protection against temperature extremes or variations; durability and resistance to degradation under conditions of use; protection from undesirable external forces such as bacterial attack, oxidation, or chemical reaction such as corrosion; and the like. Because a variety of specific chemical and/or physical properties is frequently desired for a given application, the rheology modification agent has frequently been used in conjunction with other types of agents or additives, in order to produce a final fluid suitable to a given application. However, since it is generally desirable to reduce the number of such agents or additives as much as possible, in order to facilitate the ease of production and use and, therefore, to also minimize cost, it is desirable to employ a rheology modification agent which offers the greatest number of benefits to the fluid for its intended use.

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A variety of general rheology modification agents are known and have been qualified for use in various specific applications. For example, polymeric materials such as xanthan gum, guar gum and polyacrylamides have historically been used as rheology modification agents in water-based drilling fluids, but have been found to be unstable in the presence of various salts encountered in some formations and in subsea drillsites. These materials also tend to exhibit undesirable susceptibility to oxidation and bacterial attack; to degradation when exposed to the shear forces exerted in the drilling process; and/or to thermal degradation above about 250 to 300 °C. They also have limited ability to maintain solids suspension upon elimination of shear forces such as those produced during pumping.

Showing better thermal stability are some of the non-polymeric materials, typically clays such as bentonite and attapulgit. For example, bentonite is relatively stable to temperature and offers the additional benefits of resistance oxidation and durability when exposed to high shear conditions. These mineral clays are often used with other types of agents or densifiers, such as iron oxide or barium sulfate, which enhance the ability of the fluid to resist pressures such as are encountered in subterranean excavations.

Unfortunately, the mineral clays, though historically popular, are not without their drawbacks for many applications. Fluids containing bentonite, though probably the most popular of the clay materials for drilling muds, are severely compromised in the presence of polyvalent cations, such as calcium and magnesium, frequently present in drilling formations, and may become so thick at higher temperatures under some circumstances that thinners must frequently be added. Other clay systems also suffer from undesirable reactivity and temperature degradation, and may not be adequately consistent in composition from batch to batch.

Combinations of clays and polymeric materials have also been employed, with the goal of extending the clay and thereby using less of it. Thus, the complexity of the composition is increased and therefore its cost and/or difficulty of preparation, particularly under field conditions. Typical extenders useful with bentonite systems include polyacrylamide. Unfortunately, the weaknesses of the extending polymer, such as thermal instability and the like, may then dominate the characteristics of the fluid as a whole.

In response to the above-cited problems, those skilled in the art have developed a number of newer agents based on hydrous aluminum compounds. In many cases the hydrous aluminum compounds must be formed *in situ*. This method of preparation results in formation of relatively large amounts of reaction salts which may then cause corrosion of metals such as drill bits, or may undesirably interfere with other performance additives. Furthermore, control of the reaction for *in situ* preparation may be extremely difficult, depending upon the final application, for example, in the mud pit of a drilling rig.

Also used with some success for rheology modification are crystalline layered mixed metal hydroxides, wherein Li, Mg, Cu, Zn, Mn, Fe, Co, and Ni are part of the layered structure; and also other metal aluminates. Of particular note are the "gelling" materials disclosed in U.S. Patents 4,664,843 and 4,790,954, which disclose agents that exhibit not only thickening in general, but also a variable rheology defined as "pseudoplasticity". Such rheology is characterized by an ability to flow, to a determinable extent, upon exertion of a given shear force, such as that exerted by an actively revolving drill bit or during pumping from or into the mud pit of a drilling rig. The treated fluid then returns to a significantly higher, and again determinable, viscosity when the shear force is removed. These newer rheology modification agents thus are

particularly well-suited to solids suspension combined with ease of use. However, in many cases they still suffer from some of the problems associated with the polymeric, clay and combination agents, such as limited inhibition of reactivity with some cations, undesirable toxicity, temperature limitations, insufficient lubricity, and the like. In particular, many of these agents are extremely expensive and thus impractical for drilling-scale applications in particular.

It would therefore be highly useful in the field to identify a family of agents which impart rheology modification to aqueous systems, such that their viscosity levels, with or without application of shear forces, can be optimized at each point in time according to the desired application; which exhibit desirable temperature resistance, lubricity, inhibition of reactivity, and resistance to geological formation pressure; and which are not cost-prohibitive for large scale application.

BRIEF DESCRIPTION OF DRAWINGS

FIG. 1 is an X-ray diffraction pattern, using Copper K-alpha radiation, of the Example 16 sample which has been heated as described.

FIG. 2 is an X-ray diffraction pattern of the Comparative Example E sample which has been heated as described.

FIG. 3 is an X-ray diffraction pattern of the Example 16 sample after combining with water for 0.5 minutes and then redrying.

FIG. 4 is an X-ray diffraction pattern of the Comparative Example E sample after combining with water for 0.5 minutes and then redrying.

FIG. 5 is an X-ray diffraction pattern of the Example 16 sample after combining with water for 12 minutes and then redrying.

DISCLOSURE OF INVENTION

The present invention provides such a family of agents and rheology modified aqueous compositions. It includes a method of making a rheology-modified aqueous composition comprising admixing a material or materials whose constituents substantially conform to the proportions of the empirical formula



where M' represents at least one divalent metal cation and m is an amount of from greater than zero to about 8; where M'' represents at least one trivalent metal cation and n is an amount of from greater than zero to about 6; where A is an anion or negative-valence radical that is monovalent or polyvalent, and a is an amount of A ions of valence q , provided that if A is monovalent, a is from greater than zero to about 8, and if A is polyvalent, a is from greater than zero to about 4; where B is a second anion or negative-valence radical that is monovalent or polyvalent, and where b is an amount of B ions of valence r and b is from zero to about 4; provided $(m+n)$ is greater than or equal to 1; further provided $qa + br$ cannot be greater than $2m+3n$ and provided that qa cannot equal $2m+3n$; and still further provided that $(2m+3n+qa+br)$ is less than 3; and where xH_2O represents excess waters of hydration, with x being zero or more; with at least a clay and water to form a rheology-modified aqueous composition. In certain preferred embodiments $(2m+3n+qa+br)$ is less than 2, more preferably less than 1, and most preferably less than 0.5.

The present invention further includes a method of making a rheology-modified aqueous composition comprising calcining a hydrotalcite or hydrotalcite-like compound, as defined hereinbelow, and admixing therewith at least a clay and water. The invention also comprises a dry composition useful for rheology modification of aqueous fluids

comprising a calcined hydrotalcite or hydrotalcite-like compound and a clay, and a method of making such composition.

Finally, the present invention still further includes a rheology modified aqueous composition useful for subterranean excavation comprising a calcined hydrotalcite or hydrotalcite-like compound, a clay, water and, optionally, an aluminum oxide, a nitrogen-containing compound, or both, and a method of making such composition.

Such aqueous compositions preferably exhibit a stress-dependent fluidity, as defined hereinbelow, which makes them particularly, though not solely, suitable for use as a drilling fluid, milling fluid, or mining fluid. These compositions can be prepared neat or in situ and preferably also exhibit desirable solids suspension capability; desirable inhibition, as shown by incidence of corrosivity and other reactions; low toxicity; and excellent thermal stability; when compared with other known rheology modification agents. They are also generally not prohibitively expensive for large scale applications.

MODES FOR CARRYING OUT THE INVENTION

The present invention provides a novel family of compositions which can be classified generally as dry rheology modification agents useful in water and water-based fluids of many types, and the fluids themselves as modified by the agents. Because the precise nature of the chemical structures resulting in the presence of water is subject to hypothesis, the fluid compositions are defined herein via their method of production; nonetheless, it is stressed that the resulting rheology modification cannot be explained as a result of known physical chemistry interactions, such as those which are characteristic of thixotropic, Newtonian, non-Newtonian, pseudoplastic, dilatant, Bingham plastic, or rheopexic behavior. The rheology modification exhibited by the aqueous compositions herein are therefore described simply as providing "stress-dependent fluidity", which

means that the compositions in unstressed state form an elastic solid which then is fluidized to form a true fluid (not plastic flow or plastic deformation) under certain types of stress. Flow of the fluid along or over a substrate is then measurable and can be defined by one or more of the normal descriptions of fluid flow, including plug flow, turbulent flow, laminar flow and so forth. Stresses inducing such flow are desirably primarily mechanical in nature, such as the application of shear forces exerted during drilling, mining, and milling processes, and the like, and thus the term "stress-dependent fluidity" as applied herein does not refer to fluidization which may result in some compositions due solely to increases in temperature. In fact, since many of the applications for which the compositions of the present invention may be applicable are at relatively high temperatures (preferably more than 200°F, more preferably more than 300 °F, and most preferably more than 400 °F), it is preferable that the compositions exhibit stress-dependent fluidity that is based upon mechanical stresses only, and which is also preferably not amenable to fluidization based upon simple load-bearing, such as in the column formed within a drillhole, as well as geological formation pressures.

"Stress-dependent fluidity" also requires that the phase change, from elastic solid to true fluid, be fully reversible, ensuring that the elastic solid phase of the material is reinitiated as rapidly as possible upon cessation of the mechanical stress, and that the physical shape of the elastic solid precisely imitate the physical shape of the fluid just prior to reinitiation. This ability to "gel" rapidly, using the term "gel" colloquially and without reference to the precise nature of the chemical and/or ionic bonding and/or composition of the material, is particularly important in applications such as drilling and mining, where solids suspension is critical in maintaining the integrity of the excavation during work stoppages and where pumpability must be easily reinitiated in order to ensure restarts. Those skilled in the art will understand that the term "drilling" is used

herein in its broadest meaning, to include not only the field of exploitation of geological deposits such as petroleum and/or natural gas, but also any technical accessory drilling, including but not limited to tunneling, so-called "river crossing", the sealing of dump sites, water well drilling, construction applications in general, and the like. Another way of describing the interaction of the rheology modification agent with clay in the presence of water is that the agent contributes in setting up molecular forces of some kind that initiate and carry out a physical "ordering" of the particles of clay in what is believed to be an electrostatically-maintained double-layer conformation, and at such a rapid rate, that their precipitation is prevented and they are essentially "frozen" (using the term colloquially and without reference to temperature change) into place. These molecular forces are sufficiently substantial to produce a true yield point, and thereby to support very significant weight, including columnar-applied weight, such as in a drillhole, which may include tons of soil, rock, and other debris, as well as the contributing weight of the water, clay, and agent itself.

A key starting material for the present invention is a material which conforms substantially to the empirical formula



as defined hereinabove. Alternatively, a combination of materials which can contribute the proportions of constituents of the above empirical formula can be employed. While M' can represent any divalent metal cation of the Groups IA, IIA, VIIB, VIII, IB or IIB of the Periodic Table, preferred divalent cations are Mg, Ca, Mn, Fe, Co, Ni, Cu, and Zn, and more preferred are Mg and Ca. M'' is a trivalent metal cation selected from Groups IIIA or VIII, but preferred are Al, Ga and Fe, and more preferred is Al.

There must also be present at least one anion or negative-valence radical, A, and in some cases one (or more) additional anions or negative-valence radicals, B, may also be

present. Examples of these anions and negative-valence radicals include carbonates, amines, amides, chlorides, oxides, and the like. Preferred therefor are carbonates, oxides and amides.

In selecting this first required material or combination of materials, such that the proportions of constituents substantially conform to the given empirical formula, it will be seen that it is particularly convenient to select a hydrotalcite or "hydrotalcite-like" compound or a mixture thereof. Hydrotalcite itself has the chemical formula $\text{Mg}_6\text{Al}_2(\text{CO}_3)(\text{OH})_{16} \cdot 4(\text{H}_2\text{O})$. However, varying proportions of each of these constituents (Mg, Al, and the anions CO_3 and OH) characterize related natural and synthetic minerals within the empirical formula used in the broadest claim hereinbelow, and are thus considered to be "hydrotalcite-like" compounds. Such compounds may also be selected for use in the present invention.

One possible source for such hydrotalcite and "hydrotalcite-like" compounds is, interestingly, waste streams from bauxite processing, which may include each of the constituents of the empirical formula in various forms, such as MgO , Al_2O_3 and its hydrates, $\text{Mg}(\text{OH})_2$ and its hydrates, $\text{Na}_2(\text{CO}_3)$ and its hydrates, $\text{Ca}(\text{OH})_2$, $\text{Fe}(\text{OH})_2$ and its hydrates, hydrotalcite itself and its hydrates, and various cellulosics. An example of this waste-stream material is sold by Alcoa Corporation under the tradename denomination "HTC-GL". This product, which is characterized by the manufacturer as a "green liquor", is preferably "unwashed". This means that most or substantially all of the soluble materials which are not included in the empirical formula (for example, $\text{Na}_2(\text{CO}_3)$, $\text{Ca}(\text{OH})_2$, $\text{Fe}(\text{OH})_2$, hydrates thereof, and cellulosics) are still present, along with the constituents that are included in the formula. Alternatively, a similar mixture can be prepared to mimic "unwashed HTC-GL".

Once a material source or sources for the constituents of the empirical formula has been selected, it is particularly desirable to calcine the material or materials. This calcination is heating sufficiently to, first, drive off essentially all water of hydration and, as the temperature continues to rise, to also significantly alter any crystal, i.e., "ordered", structure involving atoms of the empirical formula, but to a point less than fusion. For example, in the case of hydrotalcite, this means that the calcination is carried out to a point less than, i.e., prior to, formation of spinel, which is essentially MgAl_2O_4 which is fused and therefore highly inert. Without wishing to be bound by any one theory as to the effect of calcination, it is hypothesized that, in the case of hydrotalcite and hydrotalcite-like materials, particularly as comprised in the Alcoa "HTC-GL" ("unwashed") product, the calcination process, which occurs at an inherently basic pH, brings about a series of reactions that include dehydration, but which also go beyond simple dehydration to include changes in the crystallography and/or orientation of certain components, and possibly actually collapse of certain crystal structures such that unit sizes (a-spacings) may be significantly reduced when compared with otherwise-identical but uncalcined material. One indication of this is that the coordination state of the aluminum, as determined by nuclear magnetic resonance (magic angle spinning) (" NMR_{mas} "), changes from its pre-calcination level of essentially 100 percent octahedral, to a level which is more than 50 percent tetrahedral after calcination. Thus, the crystal attains a lower general energy state which indicates it has at least partially collapsed.

This result is in marked contrast to what has been termed in the field of minerals and inorganic chemistry as "activation" of metal hydroxides and hydrous metal oxides, sometimes in the presence of CO_2 , to produce arid, friable (easily decrepitated) small crystals of colloidal size. This "activation" process has been defined in the literature as

involving a significantly lesser degree of heating, to a point of removal of excess waters of hydration. See, for example, U.S. Patent 5,232,627 in which heating above about 700°C for hydrotalcite is proscribed on the basis that higher temperatures result in a material which is not "activated". Comparative Examples C and D hereinbelow confirm that finding and illustrate the reduced efficacy of certain hydrotalcite starting materials when heated to temperatures above about 700°C. It is hypothesized that calcination of these particular materials results in replacement of all hydroxyl constituent by an oxide constituent. Such compositions are, by definition, outside of the scope of the present invention as defined by its broadest claim, wherein it is provided that qa cannot equal $2m+3n$. U.S. Patent 4,748,139 also discloses formation of dense spinel structures at temperatures above 1000°C. Since these structures are inert, it is concluded that optimization of the compositions of the present invention occurs at a temperature in excess of 700°C, but below the point of fusion. It is interesting to note that commonly used "Loss on Ignition" ("LOI") tests illustrates that typical "activated" compositions show a significantly greater weight-percent (wt.%) decrease than do the compositions of the present invention, strongly suggesting that comparable heating of these "activated" materials simply removes waters of hydration (see Example 14 and Comparative Example A hereinbelow). Such tests are described by various American Society of Testing Materials (ASTM) specifications, such as C 25-96a and C 146-94a, which vary as to time, temperature and heating profiles, but in general the procedure can be described as comprising the steps of drying a sample; pre-igniting and weighing a crucible; placing the sample into the crucible and placing it in a hot furnace (usually 1000°C for 1 hour); cooling in a dessicator; and then weighing.

Of particular significance is the fact that $NMR_{(mas)}$ analysis shows that, at very most, from about 10 to about 15 percent of the "activated" material (and in most cases

substantially less, that is, from about 3 to about 4 percent) has been reduced to the tetrahedral aluminum coordination state via typical thermal treatment at 400 to 700°C, whereas the materials of the present invention, heated to a higher temperature and having important compositional distinctions, are shown to have reached a level of more than 50 weight percent tetrahedral, particularly interesting considering the fact that prior to such heating the materials useful in the present invention are up to 100 percent octahedral.

The calcination step employed in some embodiments of the present invention is carried out by heating to a temperature preferably from about 750°C, more preferably from about 800°C, to about 1500°C, more preferably about 1100°C. Most preferably heating is carried out at a temperature of about 900 °C to 1000°C. Such heat can be supplied by any means standard in the art, e.g., a rotary kiln, and is preferably carried out for a period of time sufficient to result in a rheology modified composition when the calcined material(s) is(are) mixed with clay and water. In most instances this requires at least about 60 minutes at about 750°C; preferably at least about 30 minutes at about 850°C; and more preferably at least about 15 minutes at about 1000°C. Time and temperature are preferably balanced to bring about the desired result, with higher temperatures obviously requiring less time. A ramping process is desirable and easily effected in a rotary kiln, with a suggested total ramp time to 1000°C of at least about 45 minutes, including a cooldown period to lower the product temperature to about 500°C.

It will be obvious to those skilled in the art that this calcination step may not be necessary in some cases, provided the constituents and proportions as defined in the empirical formula are present. See, e.g., Example 17. However, when a material such as a hydrotalcite or a hydrotalcite-like material is employed in the present invention, such calcination step is preferred to produce a rheology modification agent which is significantly more efficient at imparting the desired stress-dependent fluidity

characteristics to the clay/water combination. The calcination step is still more preferred when a combination material, such as the Alcoa HTC-GL (unwashed) or a similar mixture material, is employed. It is noteworthy that the additional constituents present in this unwashed material help to ensure conformation, after calcination, with the claim's requirements.

In alternative embodiments where a discrete combination material, such as Alcoa's HTC-GL (unwashed), is not used, such additional constituents can be added separately. Organics such as cellulose, glycols, and other organics that are capable of being combusted during calcination can be used as described in U.S. Patent 4,492,678 to Yoo et al.. Evolution of a momentary acid phase during combustion tends to increase the reaction capability of the composition with clay. Where organics are added, they are employed in an amount ranging from about 0.5 to about 10 weight percent based on final weight of the composition described by the empirical formula in Claim 1.

Sodium, especially in high concentrations of greater than about 1,000 ppm, preferably greater than about 10,000 ppm, has also been shown to improve the efficiency of the rheology modification agent with clay. Such is preferably present or employed in an amount from about 10,000 ppm to about 100,000 ppm. Interestingly, the presence of sodium in an uncalcined starting material such as a hydrotalcite operates to destroy or substantially reduce reactivity with clay, i.e., it prevents the material from being "activated", defined as being altered to enable it to absorb anions, particularly chromate, under the "CAT" test ("Colorimetric Adsorbance Test", as described in U.S. Patent 5,603,839 to Fitzgerald et al., incorporated herein by reference, measuring the amount of chromate $[\text{CrO}_2^-]$ ion which remains in solution), as it is heated to, and through, the 400 to 700°C temperature range. Without wishing to be bound by any theory, it is hypothesized that this prevention of "activation" in this temperature range is helpful in

enabling the composition to undergo the very different changes, and in particular the changes in crystallographic structure, that occur in the 750°C and higher temperature range that is employed during calcination. X-ray diffraction analysis indicates that the crystal formation progresses from ambient, at which the magnesium and aluminum may be combined in a brucite structure; to the 400 to 700°C range, where waters of hydration have been removed and MgO crystals having therein substituted aluminum form ("activation"); to the 750°C and higher range, where materials whose constituents will now conform to the formula undergo removal of the aluminum from the MgO crystals and its reformation as AlO crystals as a separate phase, the MgO and AlO aligned near each other as essentially layers of "sols", "sol" being defined as pure mono metal oxide particles having diameters from about 75 microns to about 300 microns (see, e.g., Snow, et al., *Chemical Engineers' Handbook*, 5th edition, p. 64) with the composition as a whole still adhering to the formula given in Claim 1 hereunder; to the 1300 to 1500°C range, where ceramicization occurs and a true, inert spinel forms. Despite this mechanistic theory, however, it is important to remember that, while heating certain specific materials provides one, relatively simple and inexpensive way to produce the formulaically-defined material, heating *per se* is not necessary provided that the final material adheres to the formula. Nonetheless, heating as described tends to improve efficacy.

In preparing the dry rheology modification agents of the present invention it is preferable to combine the material or materials, whose constituents substantially conform to the empirical formula, with a clay. The clay is preferably a smectitic clay of any type, with preferred clays being selected from bentonite, chlorite, polygorskite, saconite, vermiculite, halloysite, sepiolite, illite, kaolinite, attapulgite, montmorillonite, Fuller's earth, mixtures thereof, and the like. "Beneficiated" clays, which have been

chemically altered by addition of an organic polymer or a sodium compound, of any of the above types are also useful. In one embodiment the selected clay component is combined via mixing with the defined material or materials to form a dry composition which is particularly suitable for shipping and storage.

In another embodiment the dry composition is combined with water to form a rheology modified aqueous composition. This composition may itself serve as a rheology modification agent, to be added to more water to form a final, useful rheology modified aqueous composition; or the dry composition may be added to the total amount of desired water directly to form such final composition.

Additional components may also be added, to either the dry composition or to the fluid composition. Such additional components most preferably include at least an aluminum oxide, a nitrogen-based compound, or a combination thereof. These additional components preferably serve to increase the temperature resistance of the final rheology modified aqueous composition, which is particularly desirable for applications such as drilling, milling and mining. It is preferred that the final rheology modified composition have a temperature resistance based upon a desired application, but those skilled in the art will be able to balance the amount of these additional additives to achieve a given temperature stability, which is defined as the range of temperature within which the desired phase transformations, defined as stress-dependent fluidity, are not disrupted and undesired degradation of the composition as a whole does not occur. With appropriate amounts of one or both of these additives, it is possible to achieve a temperature stability of the rheology modified aqueous composition of up to at least about 300°F, more preferably to at least about 400°F, and most preferably to at least about 450°F.

It is preferred that, if an aluminum oxide is added, it is selected from either crystalline forms, such as, for example, boehmite or gybsite (also called aluminum trihydroxide), or amorphous forms, such as those exhibiting either chi or rho orientation. This aluminum oxide additive component is preferably present in an amount of from at least about 5 weight percent, more preferably from about 10 weight percent, most preferably from about 15 weight percent, to about 35 weight percent, more preferably to about 25 weight percent, based on weight of the dry composition. If the nitrogen-containing compound is employed, it is preferred that it be water-soluble and present in an amount of from about 10, more preferably from about 30, most preferably from about 35, to about 60, more preferably to about 55, weight percent based on the dry composition. The nitrogen-containing compound may be selected from any compounds of that description, including, for example, urea, propionamide, acetamide, thiourea, amine compounds, mixtures thereof, and the like. In general, materials of greater chain length and/or excessive amounts of the nitrogen-containing additive may disrupt the ability of the agent to produce the desired thickening in an aqueous system, and some materials, such as acetamide, may present toxicity problems. Urea is preferred due to its short chain length, easy availability and low cost. It is further noted that the preferred amounts described here would be preferably doubled if the given additive is measured as weight percent of just the material or materials whose constituency conforms to the given empirical formula. This means that it is preferred that the formula-defined material, which is preferably a calcined hydrotalcite or hydrotalcite-like material, preferably constitutes from about 40 to about 60, more preferably about 50, weight percent of the dry composition. It is more preferred that both the aluminum oxide and the nitrogen-containing compound be included in the dry rheology modifying composition.

It is also possible to add additional, primarily basic, compounds, such as caustic or sodium carbonate, which tend to reduce the "onset" time prior to the initial maximization of viscosity, i.e., between mixing and the formation of the elastic solid. Such addition of "onset promoters" may not be needed if the aqueous fluid is already sufficiently basic. Such may be the case when the Alcoa HTC-GL (unwashed) product or another combination of materials of relatively high pH is employed. It is preferable that the pH of the final product be from about 9 to about 14, more preferably from about 10 to about 11. However, it has been noted in the present invention that its rheological characteristics may be maintained even at very low pH, in some cases to as low as about 3, which makes it particularly desirable for use in low pH geological formations.

Other additives which may be useful in rheology modified aqueous compositions of the present invention include weighting agents, such as calcium carbonate, barium sulfate, and/or magnetite (Fe_3O_4); fluid loss control agents, such as starches and carboxymethylated starches; lubricity agents, such as glycol and glycerine; reaction inhibition agents, such as soluble potassium salts including potassium chloride and potassium acetate; and the like.

Water is also clearly necessary to form the rheology modified aqueous compositions of the present invention. Deionized or distilled water are strongly preferred, to limit the occurrence of undesirable side reactions. Its order in mixing is not critical but in some cases may affect the desired efficiency of the rheology modification. It is thus possible to combine, in dry form, the formula-defined material or materials with the clay and, optionally, other materials such as the aluminum oxide and/or nitrogen-containing compound as well as additional agents tailored to provide the desired range of properties of the final composition, and then to add the resultant dry formulation to the water; or to add either the formula-defined material or the clay to the

water first and then add the other dry constituents in any order thereafter. A particularly preferred order is to combine the clay and water first, adjust the pH to the desired range of from about 9.0 to about 10.5, preferably by addition of a base such as sodium carbonate, and then add the formula-defined material thereto. It may be desirable in some cases to limit time of exposure of the dry composition, and/or of any dry components thereof, to air due to the inherently hygroscopic nature of some of the components, for example, the hydrotalcite and hydrotalcite-like materials. For some applications, e.g., in drilling rig mud pits, it is preferred to combine the clay and water first and then add the formula-defined material and other additives thereafter. Optimization of mixing via any known mechanical means, including for example use of impeller devices, rotational mixing, or other inducement of turbulence, is desirable to ensure consistency in performance.

Regardless of mixing order, the proportions of the three main components of the rheology modified aqueous compositions are most conveniently calculated based upon their ratios and upon their weight percentage in the water or aqueous fluid as a whole. It is preferred that the ratio of clay to the formula-defined component range from about 999:1 to about 5:1, more preferably from about 99:1 to about 9:1. Thus, the clay is preferably present in any concentration which increases the viscosity of the aqueous fluid, preferably from about 0.20, more preferably from about 2.0 weight percent, preferably to about 45.0, more preferably to about 30.0, weight percent, based upon the weight of the final rheology modified aqueous composition.

In partial summary, it will be noted that, because of the variety of mixing options represented hereinabove, it is possible to prepare a fully dry composition, suitable for shipping, storage and/or later hydration; a liquid composition, particularly suited to small scale batching; or a liquid composition prepared in situ, such as would be encountered

when either the fully dry composition or small scale liquid composition is added to a much larger liquid environment, such as that encountered in a drilling rig mud pit. The final result, using any of these compositions, will be a viscosified aqueous composition which can be used at a wide variety of temperatures. Of particular importance is the fact that these compositions can be used in drilled wells having temperatures ranging from preferably about 45°F, more preferably about 70°F, to about 450°F, more preferably to about 350°F. In addition to their thermal stability and predetermined maximum viscosity, which is preferably the “gelled” elastic solid phase, they also preferably exhibit excellent stress-dependent fluidity. In general, the reduction in viscosity upon stress application, also referred to as “shear-thinning”, can be graphically predicted, with the relationship between viscosity (defined in centipoise) being substantially linear when plotted against shear rate (defined as sec^{-1} , which is a log scale). Finally, under conditions of actual use the phase transition from elastic solid to true fluid under shear conditions is rapid, preferably within about 2 minutes, more preferably effectively instantaneous, and the return to the elastic solid, or “gelled” state, occurs preferably within about 10 minutes, more preferably within about 5 minutes, and most preferably within about 0.5 minute. This last quality enables the composition to suspend drill, mill and mining solids particularly well upon cessation of shear forces such as those exerted by drill bits or during pumping. The resultant composition is furthermore preferably durable, exhibiting no or reduced reduction in its ability to make such rapid viscosity transitions upon intermittent and repeated applications of shear and in a wide variety of environments, including cation-rich environments, for uses such as subterranean excavations and a wide variety of other drilling, milling and mining applications.

These and other properties of the present invention will be further illustrated via the following examples, which are meant to be for illustrative purposes only and are not

meant to limit, nor should they be construed as limiting, the scope of the invention in any way.

Example 1

A quantity of Alcoa Corporation's "HTC-GL", described by the manufacturer as a green liquor ("unwashed") hydrotalcite slurry, is ramp-calcined in a rotary kiln to a temperature of about 1000°C, with the maximum temperature maintained for about 15 minutes and a total ramp time of about 45 minutes. About 0.43 g of this material is mixed with about 350 g of an aqueous 2.0 weight percent sodium bentonite clay dispersion. Mixing is carried out using a Hamilton Beach mixer for about 2 minutes, and then 0.15 g of sodium carbonate is added, along with a sufficient amount of a 30 percent by weight aqueous sodium hydroxide solution, to raise the fluid pH to about 10.5. The resulting composition thickens virtually immediately and shear is continued for about 10 minutes thereafter.

The composition's rheological and fluid loss properties are tested using standard methodology as described in detail in Manual of Drilling Fluids Technology, 1985, NL Baroid/NL Industries Inc., with the following results:

Yield Point	195
Plastic Viscosity	20
6 RPM* Reading	96
3 RPM Reading	72

*RPM means revolutions per minute.

The fluid loss properties of the composition are tested by placing about 300 ml of it in a standard fluid loss testing apparatus and pressurizing to 100 psig. The discharge fluid is collected for about 30 minutes and found to total greater than 40 ml.

Example 2

A composition is prepared according to Example 1, except that following its preparation the composition is heated to about 200°F for about 40 hours. The composition is then allowed to cool to about 80°F and is then tested as in Example 1, with the following results:

Yield Point	135
Plastic Viscosity	10
6 RPM Reading	80
3 RPM Reading	65

Example 3

A mixture is prepared according to Example 1, except that following its thickening, about 4 g of "Starpac II", a carboxymethylated starch fluid loss control agent sold by Chemstar Corporation, is added. The composition is then allowed to shear for the additional 10 minutes, as in Example 1, and then tested with the following results:

Yield Point	69
Plastic Viscosity	17
6 RPM Reading	68
3 RPM Reading	65

Example 4

About 0.45 g of Alcoa Corporation's "unwashed" HTC-GL product, calcined as described in Example 1, is prepared and mixed with about 0.15 g of aluminum oxide hydroxide ($\text{AlO}(\text{OH})$). The resulting admixture (about 0.60 g total) is then added to about 350 g of an aqueous 2.0 weight percent sodium bentonite clay dispersion and mixed on a Hamilton Beach mixer for about 2 minutes. About 0.15 g of sodium carbonate is then added, and also a sufficient amount of a 30 percent by weight sodium hydroxide solution to adjust the pH of the composition to about 10.5. At this point the

composition thickens almost immediately and is then allowed to shear for an additional 10 minutes. The composition's rheological properties are then tested as in Example 1 with the following results:

Yield Point	145
Plastic Viscosity	17
6 RPM Reading	65
3 RPM Reading	60

The composition's fluid loss properties are also tested as described in Example 1. The discharge fluid is collected for about thirty minutes and the total is found to be greater than about 40 ml.

Example 5

A composition is prepared according to Example 4, and the final composition is then heated to 250°F for 40 hours. The composition is then allowed to cool to about 80°F. Its rheological properties are then tested with the following results:

Yield Point	105
Plastic Viscosity	9
6 RPM Reading	28
3 RPM Reading	25

Example 6

A composition is prepared according to Example 1, except that about 0.45 g of the calcined ("unwashed") HTC-GL is mixed with about 0.215 g of aluminum oxide hydroxide are used. Additionally, about 0.215 g of urea is mixed with the aluminum oxide hydroxide prior to incorporation with the aqueous sodium bentonite clay dispersion. Testing upon completion of the composition yields the following results:

Yield Point	175
Plastic Viscosity	18
6 RPM Reading	78
3 RPM Reading	72

Example 7

A composition is prepared as in Example 6. Following completion of preparation it is then heated to about 400°F for about 40 hours, and is then allowed to cool to about 80°F.

Tested yields the following results:

Yield Point	74
Plastic Viscosity	9
6 RPM Reading	54
3 RPM Reading	50

Example 8

A drilling mud is prepared by mixing about 0.43 lb of calcined HTC-GL ("unwashed") material, calcined as described in Example 1; 7 lb of sodium bentonite clay; and deionized water to make 1 barrel. About 0.15 lb of a mixture of sodium carbonate and 0.04 lb of a 30 weight percent sodium hydroxide solution are then added to adjust pH to about 10.5 and the composition is stirred for about 20 minutes. The composition is then allowed to set for about 10 minutes. It is then heated to 200°F for about 40 hours. At about the sixteenth hour of the heating, it is mixed for about 1 minute, allowed to set for about 5 minutes, and then rheology-tested. In the table below, the results are shown and compared with the results at 0 hours (from Example 1) and 40 hours (from Example 2) for compositions prepared using the same proportions and methodology.

	0 hrs. (Ex. 1)	Time 16 hrs.	40 hrs. (Ex. 2)
Yield Point	195	194	135
Plastic Viscosity	20	11	10
6 RPM Reading	96	115	80
3 RPM Reading	72	90	65

Example 9

The procedure of Example 8 is followed except that the heating is done at 250°F. The results are as follows:

		Time	
	0 hrs.	16 hrs.	40 hrs.
Yield Point	176	26	12
Plastic Viscosity	27	14	13
6 RPM Reading	80	4	3
3 RPM Reading	65	2	2

Example 10

A composition is prepared as in Example 8, except that "Starpac II - grade P-3228", a carboxymethylated starch fluid loss prevention additive sold by Chemstar Corporation, is also added following the 20-minute stirring. The composition is then mixed for an additional 15 minutes. The final composition is allowed to set for 10 minutes and then a portion is heated for 40 hours at 200°F and, following the procedure used in Example 8, rheology is tested at 16 hours and at 40 hours. In the table below the results are compared with the results of testing of a very similar composition as prepared in Example 3, except that a slightly different fluid loss prevention additive is used in that Example.

		Time	
	0 hrs. (Ex. 3)	16 hrs.	40 hrs.
Yield Point	69	81	70
Plastic Viscosity	17	31	36
6 RPM Reading	68	76	60
3 RPM Reading	65	73	58

Example 11

Another portion of the composition prepared for Example 10 is heated for 40 hours, but at 250°F, rather than at 200°F as in that Example. Testing results for this portion are as follows:

	Time		
	0 hrs.	16 hrs.	40 hrs.
Yield Point	64	53	46
Plastic Viscosity	16	31	27
6 RPM Reading	67	32	30
3 RPM Reading	63	28	27

Example 12

A drilling mud is prepared as in Example 8, using about 0.45 lb of calcined HTC-GL ("unwashed") material and 0.15 g of aluminum oxide hydroxide. Proportions of clay, sodium carbonate and sodium hydroxide are the same as in that Example. Following stirring for 20 minutes, the mud is allowed to set for 10 minutes and then rheology is tested. The Example shows test results for portions of mud heated at three different temperatures (200°F, 250°F and 275°F) and procedures immediately preceding testing at each step are as described in Example 8.

	Time at 200°F		
	0 hrs.	16 hrs.	40 hrs.
Yield Point	145	155	145
Plastic Viscosity	17	12	16
6 RPM Reading	65	70	60
3 RPM Reading	60	65	50

	Time at 250°F		
	0 hrs.	16 hrs.	40 hrs.
Yield Point	152	117	105
Plastic Viscosity	19	8	9
6 RPM Reading	70	40	28
3 RPM Reading	64	35	25

	Time at 275°F		
	0 hrs.	16 hrs.	40 hrs.
Yield Point	155	32	28
Plastic Viscosity	15	13	12
6 RPM Reading	75	25	18
3 RPM Reading	70	22	17

Example 13

A drilling mud is prepared as in Example 8, but using about 0.25 lb per barrel (ppb) of calcined HTC-GL (unwashed) material, about 5.5 ppb of sodium bentonite clay, and about 0.15 ppb of sodium carbonate. The pH is adjusted as in that Example, using a 30 percent aqueous sodium hydroxide solution, and the methodology of preparation is the same. Shear rate and viscosity are measured simultaneously to determine the shear thinning, or stress-dependent fluidity, characteristics of the composition, and recorded in the table below.

RPM	sec-1	dial reading	viscosity multiplier	cps
1200	2040	84	0.25	21
600	1020	84	0.5	42
300	510	78	1	78
100	170	71	3	210
6	10.2	58	50	2,900
3	5.1	57	100	5,700
1.5	2.55	56	200	11,200

When the table results are plotted as a graph in which the vertical axis represents viscosity in centipoise (cps) and the horizontal axis represents shear rate in sec^{-1} , a log scale, and in which units increase in direction away from the origin and are each equidistant and multiplied by a factor of 10 over the preceding unit, the relationship is shown to be substantially linear.

Example 14 and Comparative Example A

A composition ("Example 14") is prepared according to Example 1, then allowed to cool to about 80°F and subjected to a battery of tests for its chemical and physical properties as well as reactions with an aqueous clay (sodium bentonite). A comparative composition, obtained from Alcoa Corporation under the denomination "AHTC" ("Comparative Example A") is subjected to identical testing. The results are placed alongside each other in the following chart for ease of comparison.

	<u>Comparative Example A</u>	<u>Example 14</u>
<u>Physical Properties</u>		
Surface area, m ² /g ¹	150-300	40-60
% increase in surface area with	200-1000	10-20
Average particle size ²	0.1-0.5	20-70
Solubility, wt. %	<0.005	5-12
Loss on ignition at 1000°C, wt. %	10-25	3.5-6.5

Chemical Properties

Dual metal crystals, wt. %	95	<2
Purity (MgAl), wt. %	99.9	87-93
Heat of dispersion, btu/lb	200-300	<15
Chromium absorption, wt. % ³	>85	0
Liquid (H ₂ O) hydration rate, hrs	<0.1	>24
Gas absorption (CO ₂ /H ₂ O) (wt. gain)	>70	<10
Aluminum coordination state (Td) wt% ⁴	<25	>50
Aluminum to hydroxyl molar ratio	1:≥3	1:≤0.5
Sodium concentration, ppm	<600	>70,000

Reaction with Aqueous Clay

Critical flocculation value ⁵	Yes	No
Aggregation value	Broad	Limited
Initial thickening time, minutes:		
With shear	4-6	<0.5
Without shear	>20	2-5
Adduct reaction signature (pH change)	0.5	None
Air interference (4 days)	Severe	Slight
Rheology stability:		
pH range	4.5-14	2-14
Temperature, °C	275	<220
Potassium ion addition	Poor	Good
Clay reaction following water dispersal		
Filtrate	None	Thickens
Residue	Slight	Thickens

¹ As described in U.S. Patent 5,068,095 using the BET/N₂ process. "BET" is Brunauer-Emmett-Teller.

² As defined in Snow, et al., *Chemical Engineers' Handbook*, chap. 8, p. 7 (in solution).

³ Also called "CAT-307" test. As described in U.S. Patent 5,603,839. Suppliers of "activated hydrotalcite" use this test to indicate whether or not a hydrotalcite is "activated". To be classified as an "activated hydrotalcite" ("AHTC"), the product must remove >75% of the anion, generally chromate, in a given time (30 minutes).

⁴ "See "Hydrotalcite decomposition mechanism" in *J. Phys. Chem.*, vol. 100, 8535-8542 (1996).

⁵ As described in Gray, et al., *Composition and Properties of Oil Well Drilling Fluids*, 4th ed., chap. 4, pp. 138-180.

Comparative Example B

About 7.0 g of sodium bentonite clay is mixed with 343 g of water to make a 2.0 weight percent aqueous dispersion. Mixing is carried out using a Hamilton Beach mixer for about 60 minutes, and then 0.15 g of sodium carbonate is added, along with a sufficient amount of a 30 percent by weight aqueous sodium hydroxide solution, to raise the fluid pH to about 10.5. The resulting composition is sheared for an additional 30 minutes.

The composition's rheological and fluid loss properties are tested using standard methodology as described in the previous Examples, with the following results:

Yield Point	4
Plastic Viscosity	4
6 RPM Reading	0
3 RPM Reading	0

The fluid loss properties of the composition are tested by placing about 300 ml in a standard fluid loss testing apparatus and pressurizing to 100 psig. The discharge fluid is collected for about 30 minutes and found to total 22 ml.

Comparative Example C

A quantity of Alcoa Corporation's "washed" HTC-GL green liquor hydrotalcite slurry is ramp-calcined in a muffle kiln to a temperature of about 1000°C, with the maximum temperature maintained for about 15 minutes and a total ramp time of about 45 minutes. About 0.43 g of this material is mixed with about 350 g of an aqueous 2.0 weight percent sodium bentonite clay dispersion. Mixing is carried out using a Hamilton Beach mixer for about 2 minutes, and then 0.1 g of sodium carbonate is added, along with a sufficient amount of a 30 percent by weight aqueous sodium hydroxide solution, to raise the fluid pH to about 10.5. The resulting composition is sheared for an

additional 30 minutes. Testing is carried out as described hereinabove, with the following results:

Yield Point	5
Plastic Viscosity	4
6 RPM Reading	1
3 RPM Reading	0

The fluid loss properties of the composition are tested by placing about 300 ml of it in a standard fluid loss testing apparatus and pressurizing to 100 psig. The discharge fluid is collected for about 30 minutes and found to total 24 ml.

Comparative Example D

A quantity of Alcoa Corporation's "HTC-24", described by the manufacturer as a spray-dried, "washed" hydrotalcite powder, is ramp-calcined in a muffle kiln to a temperature of about 1000°C, with the maximum temperature maintained for about 15 minutes and a total ramp time of about 45 minutes. About 0.43 g of this material is mixed with about 350 g of an aqueous 2.0 weight percent sodium bentonite clay dispersion. Mixing is carried out using a Hamilton Beach mixer for about 2 minutes, and then 0.15 g of sodium carbonate is added, along with a sufficient amount of a 30 percent by weight aqueous sodium hydroxide solution, to raise the fluid pH to about 10.5. The resulting composition is sheared for an additional 30 minutes.

The composition's rheological and fluid loss properties are tested as described in previous Examples, with the following results:

Yield Point	8
Plastic Viscosity	6
6 RPM Reading	1
3 RPM Reading	1

The fluid loss properties of the composition are tested by placing about 300 ml of it in a standard fluid loss testing apparatus and pressurizing to 100 psig. The discharge fluid is collected for about 230 minutes and found to total 22 ml.

Example 16 and Comparative Example E

X-ray diffraction testing is performed on samples using Copper K-alpha radiation to compare the two. The X-ray diffraction pattern of the Example 16 sample is shown in FIG. 1, and the X-ray diffraction pattern of the Comparative Example E sample is shown in FIG. 2. It is noted that the X-ray diffraction pattern of the Example 16 sample (FIG. 1), which is Alcoa's "HTC-GL (unwashed)" which has been calcined at 850-900°C for one about one hour to achieve a composition corresponding to the formula of the present invention, shows that, while both aluminum and magnesium are present in separate crystalline oxide form, there is essentially no aluminum substituted into the magnesium oxide crystals. In contrast, the pattern of the Comparative Example E sample (FIG. 2), which is a hydrotalcite (Alcoa's "HTC-24") which has been "activated" via heating to 450-550°C, shows significant substitution of aluminum into the magnesium oxide crystals (signified by the offset of the peaks at 42-43 and 62-63 degrees 2-theta, i.e., the magnesium oxide peaks).

The two samples are then combined with water, dried and new X-ray diffraction patterns are prepared. FIG. 3 shows the pattern corresponding to Example 16, which corresponds with that of FIG. 1, that is, there is essentially no transformation of the pattern after the addition of water and exposure to an aqueous environment for 0.5 minutes. In contrast, the X-ray diffraction pattern of the Comparative Example E sample, after combination with water followed by drying, shows very significant transformation, including those typical of meixnerite, a hydrotalcite-like mixed metal hydroxide material. A further X-ray diffraction pattern is taken of the Example 16 material, after exposing it to the aqueous environment for 12 minutes followed by redrying. Again, no significant transformation, when compared with FIG. 1, occurs. These comparisons illustrate the importance of selecting for use in the present invention

a material which falls within the scope of the given formula rather than one which has similar constituents but proportions which do not meet the formula's limitations.

Example 17

About 350 g of a 2 weight percent aqueous mixture of clay (montmorillonite) is placed in a Hamilton Beach mixer and mixed thoroughly at low speed. To this clay dispersion is added about 0.9 g of a physical mixture consisting of urea (22 weight percent), $Mg(OH)_2$ (54 weight percent), $AlO(OH)$ (30 weight percent), and $NaAlO_2$ (4 weight percent). The pH of the resulting fluid is adjusted to 11.0 using NaOH and then mixed for 20 minutes. The rheology properties are tested with the following results.

Yield Point	132
Plastic Viscosity	15
6 RPM Reading	68
3 RPM Reading	62

The fluid is then heated to 250°F and hot-rolled for 16 hours, followed by mixing for one minute and pH adjustment to 11.0. The rheology properties are then retested with the following results.

Yield Point	70
Plastic Viscosity	7
6 RPM Reading	52
3 RPM Reading	47

Example 18

A fluid is prepared as described in the first paragraph of Example 17, except that after mixing 20 minutes at pH 11.0, about 4 g of a fluid loss additive ("Starpac II – grade P-3228", a carboxymethylated starch fluid loss control agent sold by Chemstar Corporation) is added and mixed for an additional 5 minutes. Rheology properties are then tested with the following results.

Yield Point	128
Plastic Viscosity	23
6 RPM Reading	78
3 RPM Reading	69

The fluid is then heated to 250°F and hot-rolled for 16 hours. It is allowed to cool and then the rheology properties are retested with the following results.

Yield Point	74
Plastic Viscosity	17
6 RPM Reading	64
3 RPM Reading	57

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